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**Process for the production of a catalyst preparation and  
use of the catalyst preparation**

5 The invention relates to a process for the production of  
a catalyst preparation and the use thereof.

10 Solid catalysts are frequently utilised in the  
production of polyesters. Such catalysts may, for  
example, be  $\text{TiO}_2$  or compounds derived from  $\text{TiO}_2$ . For  
example, EP 0736560 describes catalysts comprising  
alkali metal titanate or alkaline earth metal titanate.

15 In heterogeneous catalysis the catalyst is not dissolved  
in the reaction medium but is present as a separate  
solid phase. In one variant on this catalysis, the  
catalyst remains as a separate solid phase within the  
system. In some cases, such as, for example, in  
catalysis of esterification processes, the catalyst is  
20 added as a powder which, however, dissolves in the  
course of the reaction. Here, a continuous transition  
from heterogeneous to homogeneous catalysis takes place.

25 The distribution of the catalyst in the reaction medium  
greatly influences the course of the reaction.

It is in principle possible to supply a heterogeneous  
catalyst in powder form to the reaction medium.  
However, in continuous processes specifically, in  
addition to the disadvantage of poor meterability of the  
30 powders, this method has deficiencies in terms of  
distributing the catalyst rapidly in the reaction  
medium. It is therefore prior art practice to utilise  
slurries of the catalyst in liquids.

35 Since the catalyst particles settle readily when the  
catalyst slurries are stored, such slurries must be  
agitated before use, because otherwise the metering of  
the catalyst would in turn not be sufficiently precise.

Stokes' law states that the rate of sedimentation  $v$  is proportional to the square of the diameters  $d$  of the catalyst spherules and to the density difference  $\Delta\rho$  between solid and liquid, and inversely proportional to the viscosity  $\eta$  of the suspending liquid:

$$v = \frac{d^2 \cdot \Delta\rho \cdot g}{18\eta} \quad (\text{Stokes' law})]$$

10 It is apparent from the equation that over time all the catalyst particles must settle, provided that their rate of sedimentation is greater than their thermal (Brownian) motion.

15 In addition to gravity, the van der Waals force of attraction, also known as the "dispersion force", also acts on the particles. If not opposed by stabilising forces, for example a like electrostatic charge, or adsorbed polymers, the dispersion force causes  
20 flocculation to take place among the particles. If the particles flocculate, then larger agglomerates arise which settle all the more rapidly in accordance with Stokes' equation. Flocculated, settled systems, on the other hand, have the advantage that they can be agitated  
25 readily.

The opposite is the case with slurries, in which sedimentation out of a flocculation-resistant distribution takes place because of the size of the  
30 catalyst particles. Apart from the disadvantage of settling to a very densely packed sediment now barely able to be agitated, a further disadvantage of such systems is that they generally contain undesirable auxiliary substances (polymeric wetting and dispersing  
35 agents).

The object of the invention is to provide a process for the production of a catalyst preparation, in which the disadvantages of the prior art, namely in particular

- 5    - excessively coarse distribution of the catalyst
- poor meterability of the catalyst powder
- poor meterability of the catalyst when settled slurries are used
- sedimentation of the catalyst to a solid bottom
- 10    sediment which cannot be agitated

are avoided, and in which the catalyst obtained is, on the one hand, sufficiently stable in storage and, on the other, can deploy the maximum catalytic activity.

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The object is achieved by a process for the production of a catalyst preparation, in which the catalyst, containing at least one inorganic compound which is solid under standard conditions, is comminuted by means of a dispersion unit into particles having a maximum average particle size  $d_{50,3}$  of 2  $\mu\text{m}$ , preferably a maximum of 1  $\mu\text{m}$ , implemented in accordance with DIN 66141 and DIN 66144, and is distributed at a concentration of from 1 to 50 wt.%, preferably 20 to 40 wt.% (relative to the finished catalyst preparation) in a liquid. The particle size is preferably measured by laser diffraction.

Surprisingly, without the aid of wetting or dispersing agents or otherwise any auxiliary substance which stabilises the suspension, the catalyst can in this way be produced as a paste which is solidified in the resting state. The paste does not begin to flow until a minimum shear stress, the yield point  $\tau_0$ , is reached.

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The yield point of the catalyst paste rises as the catalyst concentration increases and the particle size decreases, respectively. The yield point of the

- catalyst paste can be adjusted by varying the catalyst concentration and the particle size. The catalyst paste preferably has a yield point  $\tau_0$  of at least 0.1 Pa at 23 °C (room temperature), particularly preferably 1 to 30 Pa and most particularly preferably 5 to 30 Pa, at 23 °C. The yield point is determined in accordance with DIN 53019 in the Z2 DIN cylinder measuring system (described in the data sheet "Data for the standard measuring system in accordance with DIN 53019" for Viscolab/Rheolab MC 10 from Physika Meßtechnik GmbH, Stuttgart) and evaluation in accordance with Bingham. For this purpose, the sample (catalyst paste) is charged into the Z2 standard measuring system where it is first left to rest at room temperature for a period of 24 hours. Measuring takes place with 100 measuring points with a measuring time of 5 sec per measuring point within the shear rate range 0 to 100 1/s (linear ramp).
- The process described guarantees that the stable catalyst paste can be brought by agitation to a form in which it is readily meterable and can be mixed effectively into the reaction mixture, with the catalyst activity being put to optimal use. Here, the catalyst is present in colloidal form and is unusually stable in storage. Even when the suspension is stored for a year, the catalyst does not settle and retains its full activity. It is furthermore advantageous that the utilisation of auxiliary substances for creating a yield point (for example polyurethane thickeners, polyacrylate thickeners, cellulose ether, highly disperse silica) is superfluous.

- The inorganic compound which is solid under standard conditions (catalyst or catalyst constituent) can, for example, be selected from the following substances: titanium dioxide, titanium-dioxide-containing substances, titanates, zeolites, aluminium oxide, boron

oxides, germanium dioxide, antimony(III) oxide, cerium oxides, barium sulfate, zinc sulfide, silicon dioxide or mixtures of these substances.

- 5 Titanium dioxide and the titanium-containing catalysts described in EP 0736560: hydrated titanium dioxide corresponding to the composition

10  $y \text{ TiO}_2 \cdot z \text{ H}_2\text{O}$  (where  $y = 1$ ,  $z = 0.01$  to  $2$ ), or a titanate corresponding to the composition

$(\text{Me}_n\text{O})_x \cdot (\text{TiO}_2)_y \cdot (\text{H}_2\text{O})_z$  (where  $\text{Me} = \text{Li}, \text{Na}, \text{K}, \text{Rb}, \text{Cs}, \text{Mg}, \text{Ca}, \text{Sr}, \text{Ba}$ ;  $n = 1$  for  $\text{Me} =$  alkaline earth metal and  $n = 2$  for  $\text{Me} =$  alkali metal;  $x = 0.0001$  to  $6$ ;  $y = 1$ ;  $z =$   
15  $0.01$  to  $2$ ).

are particularly preferred here.

20 Stirred ball mills, ultrasonic homogenisers, or ultrasonic disintegrators, high-pressure homogenisers, dispersing equipment based on the high-power pulse-type technique, dispersing equipment based on the impact jet process (for example counter-jet-type mills) or impact stream-type mills (for example microjet dispersers) can,  
25 for example, be utilised as the dispersion unit. Stirred ball mills are particularly preferred.

The catalyst can be pre-dispersed or formed into a slurry in the liquid before the comminution. For this,  
30 dissolvers, stirrers (for example straight-arm paddle agitators, cage-type agitators or toothed-disk stirrers), Hoesch turbines or Ystral mills can be used. The catalyst is then brought to the desired particle size in the liquid by means of one of the dispersion  
35 units described above.

The catalyst can also be dry-comminuted and then formed into a slurry in the liquid in a dissolver, stirrer (for

example straight-arm paddle agitator, cage-type agitator or toothed-disk stirrer), a turbine or an Ystral mill.

5 The following relationship generally applies to the choice of particle size and catalyst concentration: The smaller the particle size, the lower the selected catalyst concentration should be within the limits indicated, and vice versa.

10 The following substances are preferably utilised individually or in mixture as the liquid: water, alcohols having 1 to 20 C atoms (for example methanol, ethanol, propanol (all isomers), butanol (all isomers), diols such as alkane diols or cycloalkane diols having 2  
15 to 12 C atoms (for example ethylene glycol, 1,3-propanediol, 1,4-butanediol, 1,4-cyclohexane dimethanol), carboxylic acids (for example formic acid, ethanoic acid, propanoic acid, butanoic acid, pentanoic acid) or fatty acids.

20 The process according to the invention can, for example, be carried out as follows (with no limitations being, however, associated therewith):

25 Where it is not already present in powder form, the catalyst material is powdered in known manner, stirred into the liquid and pre-dispersed (for example using a toothed-disk dissolver). The suspension obtained is then treated by means of a dispersion unit (for example  
30 a bead mill) until the average particle size  $d_{50,3}$  reaches a maximum value of 2  $\mu\text{m}$ , preferably a maximum of 1  $\mu\text{m}$ . The catalyst preparations produced according to the invention are used, *inter alia*, in condensation and polycondensation reactions (for example the amidation of  
35 carboxylic acids, esterifications of carboxylic acids and their hydrolysis), in transesterifications of esters, in transamidations of amides, in rearrangements

(for example alpha-pinene into camphene, aldol reaction) and in olefin polymerisation.

The catalyst preparation produced according to the invention can furthermore be used in photocatalysis. Thus, for example a  $\text{TiO}_2$ -containing catalyst preparation can be introduced into a medium which is polluted with noxious bacterial or chemical substances, with the noxious substances being oxidised in the presence of light. Areas for utilising such photocatalysts are paints, lacquers, finishes, paper, wallcoverings, self-cleaning surfaces or antibacterial surfaces.

A further possible use for the catalyst preparation is to boost the effect of catalyst enzyme systems. For this, for example, enzymes can be added during the production of the catalyst preparation to the inorganic compound distributed in the liquid. The enzymes settle out on the surface of the solid. The enzymes are advantageously present there in distributed form and are at the same time immobile.

The invention is explained in greater detail with reference to the Embodiment Example which follows:

**Example 1: production of a stable titanium-containing catalyst preparation**

280 g of a pulverulent Ti-containing catalyst having a Ti content of 48.1 wt.%, a specific BET surface of  $71.5 \text{ m}^2/\text{g}$  and an average particle size  $d_{50,3}$  of  $2.7 \mu\text{m}$ , which has been produced in accordance with the method described in EP 0736560, were stirred into 520 g monoethylene glycol by means of a toothed-disk dissolver (diameter 4 cm, speed 1300 r.p.m.) and dispersed for 45 minutes. The suspension thus formed was ground in a stirred ball mill (Dispermat SL, 125 ml grinding compartment, 100 ml glass beads, diameter 1 mm) for

20 minutes at 3000 r.p.m. and 50 minutes at 4000 r.p.m. in a circulating process. The suspension was separated from the beads and the coarse fraction by way of a sieve (56  $\mu\text{m}$  mesh). The particle size determination (laser diffraction, Helos) showed an average particle diameter  $d_{50.3}$  of 0.92  $\mu\text{m}$  with a breadth of distribution of  $B_{90/10} = 1.5 \mu\text{m}$ . The solids content was around 35 wt.%. After three days' storage at 80 °C the suspension also showed no bottom sediment formation. The dispersed suspension was charged into a Z2 DIN measuring system and left to stand at room temperature for 48 hours. Measurement took place with 100 measuring points with a measuring time of 5 sec per measuring point within the shear rate range of 0 to 100 1/s (linear ramp). The shear stress (on commencing shear stressing) of the suspension was determined at  $\tau_0 = 22.6 \text{ Pa}$  (measurement in accordance with DIN 53019, Z2 DIN, evaluation in accordance with Bingham).